

Name of the Student: _____

Max. Marks : 26 Marks

Time : 26 Minutes

Mark Schemes

Q1.

- (a) balance / scales

1

- (b)
- $\text{density} = \frac{\text{mass}}{\text{volume}}$

or

$$\rho = \frac{m}{V}$$

1

- (c)
- $0.68 = \frac{85}{V}$

1

$$V = \frac{85}{0.68}$$

1

$$V = 125 \text{ (cm}^3\text{)}$$

1

- (d) repeat readings (of volume) need taking (of each fruit) to show that the readings are close together

allow 'the same' for 'close together'

1

[6]**Q2.**

- (a) 0(.0) to 12(.0)

allow 2(.0) to 12(.0) (N)

1

- (b) mass of gas (in the syringe)

or

temperature (of the gas)

1

- (c) constant =
- 60×45

or

constant = 2700

1

$$2700 = p \times 40$$

1

$$p = \frac{2700}{40}$$

1

$$p = 67.5 \text{ (kPa)}$$

allow 68 (kPa)

1

- (d) there is more time between collisions of particles and the walls of the syringe
or
 there are less frequent collisions between the particles and the walls of the syringe

1

(causing) a lower (average) force on the walls of the syringe

1

(and) pressure is the total force per unit area

1

[9]

Q3.

- (a) the gradient for ice is steeper than the gradient for water (liquid)
allow the temperature of the ice increased faster than the temperature of the water

1

which means that less energy is needed to increase the temperature by a fixed amount

1

- (b) water took more time to vaporise than the ice took to melt

1

which means that less energy is needed to change the state from solid to liquid (than from liquid to vapour)

1

- (c) any **two** from:

- ice/water would take more time to increase in temperature
allow gradients would be less steep
- ice/water would take more time to change state
- the change in temperature with time would not be linear
allow horizontal lines would be longer

2

- (d) $E = 69\,000 \text{ (J)}$

1

$$69\,000 = 0.030 \times L$$

allow a correct substitution of an incorrectly/not converted value of E

1

$$L = \frac{69\,000}{0.030}$$

allow a correct rearrangement using an incorrectly/not

converted value of E

1

$$L = 2\,300\,000$$

or

$$L = 2.3 \times 10^6$$

*allow a correct calculation using an incorrectly/not
converted value of E*

1

J/kg

*allow a unit consistent with their numerical answer
eg 2300 kJ/kg*

1

[11]